Complete the following reactions. Circle the most favored products.

$$2. \qquad \qquad + \text{ HCI } \qquad \boxed{\text{[ZnCl}_2]} \\ \bullet$$

$$3. \begin{array}{c} OH \\ + HCI \end{array} \begin{array}{c} [ZnCl_2] \\ \end{array}$$

4.
$$\stackrel{\mathsf{OH}}{\longleftarrow}$$
 $\stackrel{[-\mathsf{H}_2\mathsf{O}]}{\longrightarrow}$

5. —OH + OH
$$[-H_2O]$$

$$6. \qquad \qquad \underbrace{ \text{OH} \qquad \xrightarrow{\left[K_{2} Cr_{2} O_{7} / H_{2} SO_{4}\right]}}_{\Delta} \\$$

$$9. \quad \xrightarrow[K_2 \text{Cr}_2 \text{O}_7/\text{H}_2 \text{SO}_4]} \xrightarrow[\Delta]{\text{OH}}$$

10.
$$\downarrow$$
 OH + \downarrow OH $\underline{[H_2O]}$

14.
$$\bigcirc$$
OH + \bigcirc OH $\underbrace{[H_2O]}$

20.
$$+$$
 HCl $[ZnCl_2]$

$$21.$$
 + NaOH \rightarrow

$$23. \qquad \text{OH} \qquad \xrightarrow{\frac{[K_2 Cr_2 O_7/H_2 SO_4]}{\Delta}}$$

$$24. \qquad \underbrace{\frac{\left[\text{K}_2\text{Cr}_2\text{O}_7/\text{H}_2\text{SO}_4 \right]}{\Delta}}$$

$$25.$$
 $\stackrel{\text{CI}}{\longleftarrow}$ + NaOH \longrightarrow

Question 2:
$$+ H_2O$$
 Rate: Very Slow

Question 3:
$$+ H_2O$$
 Rate: Fast

Question 7:
$$+ H_2O$$

Question 8:
$$+ H_2O$$

Question 9: No Reaction

Question 10:
$$+ H_2O$$

Question 11:
$$+ H_2O$$

Question 12:
$$+ H_2O$$
 Rate: Fast

Question 13:
$$+ H_2O$$
 Rate: Fast

Question 14:
$$+ H_2O$$

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Question 16:
$$+ H_2O$$
 Rate: Slow

Question 17:
$$+ H_2O$$

Question 18:
$$+ H_2O$$
 Rate: Very Slow

Question 19:
$$+ H_2O$$

Question 20:
$$+ H_2O$$
 Rate: Very Slow

Question 22:
$$+ H_2O$$
 Rate: Fast

Question 24: No Reaction