Complete the following reactions. Circle the most favored products.

1.
$$+$$
 HCl $\underline{[ZnCl_2]}$

$$2$$
. \bigcirc CI $+$ NaOH \longrightarrow

4.
$$\downarrow$$
 OH + OH $\underline{\text{[-H_2O]}}$

$$5. \qquad \stackrel{\mathsf{OH}}{\longleftarrow} \qquad \stackrel{[\mathfrak{0}]}{\longleftarrow}$$

$$6. \qquad \stackrel{\mathsf{OH}}{\longleftarrow} \qquad \stackrel{[\mathfrak{0}]}{\longrightarrow} \qquad$$

$$8. \qquad \xrightarrow{[K_2 \operatorname{Cr}_2 \operatorname{O}_7/\operatorname{H}_2 \operatorname{SO}_4]} \delta$$

11. OH + OH [
$$H_2O$$
]

14. —OH +
$$\bigcirc$$
OH [-H₂O]

$$16. \hspace{1cm} \begin{array}{c} \text{OH} \\ \end{array} \hspace{1cm} + \hspace{1cm} \text{HCI} \hspace{1cm} \begin{array}{c} \text{[ZnCl}_2] \\ \end{array} \hspace{1cm}$$

$$17. \qquad \underset{\Delta}{\text{OH}} \qquad \underset{\Delta}{\underbrace{|K_2 \text{Cr}_2 \text{O}_7 / \text{H}_2 \text{SO}_4|}},$$

18. OH +
$$\bigcirc$$
 [H₂O]

19. OH + OH [
$$H_2O$$
]

$$20. \qquad \qquad \underset{\Delta}{\text{OH}} \xrightarrow{ [K_2 \text{Cr}_2 \text{O}_7/\text{H}_2 \text{SO}_4]}$$

21.
$$+$$
 HCI $[ZnCl_2]$

$$23. \quad \stackrel{\text{OH}}{\underbrace{\qquad \qquad }_{\left[K_{2}Cr_{2}O_{7}/H_{2}SO_{4}\right]}}$$

Question 4:
$$-$$
 + H_2O

Question 7:
$$+ H_2O$$

Question 9:
$$+ H_2O$$
 Rate: Very Slow

Question 10:
$$+ H_2O$$

Question 13:
$$+ H_2O$$

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Question 16:
$$+ H_2O$$
 Rate: Slow

Question 18:
$$+ H_2C$$

Question 21:
$$+ H_2O$$
 Rate: Very Slow

 ${\color{red} {\rm Question}} \ \ 23{:} \ \ \ {\color{red} {\rm No}} \ \ {\color{red} {\rm Reaction}}$

Question 24:
$$+ H_2O$$